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Practice balancing chemical equations by changing the coefficients of reactants and products. As the equation is manipulated, the amount of each

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element is shown as individual atoms, histograms, or numerically. Molar masses of reactants and products can also be calculated and balanced to demonstrate conservation of mass.

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like earning appointments, giving directions towards clinic and answering a variety of other routine issues from patients.

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This process is summarized by a

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chemical equation. In the Balancing Chemical Equations Gizmo, look at the floating molecules below the initial reaction:  $\text{H}_2 + \text{O}_2 \Rightarrow \text{H}_2\text{O}$ .

1. How many atoms are in a hydrogen molecule ( $\text{H}_2$ )? 2
2. How many atoms are in an

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oxygen molecule ( $O_2$ )? 2 3. How many hydrogen and oxygen atoms are in a water molecule ( $H_2O$ )? H: 2 O: 1 4. In general, what does a subscript (such as the “2” in  $H_2$ )

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Check your answers for the first three formulas using the Gizmo.

$\text{AgCl}_3\text{Cu}_2$ . Ag:   1   Cl:   3  .

Cu:   2  .  $\text{Ba}(\text{AsO}_4)_2$  Ba:   1  

As:   2  . O:   8  . 2018.

$(\text{NH}_4)_3\text{PO}_4$  N:   3   H:  12   P:

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Balancing Chemical Equations Worksheet 1 - ANSWERS  
1.  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$   
2.  $2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$   
3.  $\text{N}_2 + \text{O}_2 \rightarrow 2\text{NO}$   
4.  $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$   
5.  $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$

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## **BalancingChemEquationsSE.docx**

If you also get perplexed in balancing chemical equations, follow the tips for correct balancing chemical equations worksheet answers. Tip # 1: When you are trying to balance

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the chemical equations, you should remember that you can only change the value of coefficient in front of the element or compound, and not the subscript.

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Student Exploration: Balancing Chemical Equations Learning Objectives Students will&mldr; Adjust coefficients to balance a chemical equation. Classify chemical equations as combination (synthesis), decomposition, single



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replacement, or double replacement. Vocabulary: coefficient, combination, compound, decomposition, double replacement, element, molecule, product, reactant, single replacement ...

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