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Fraction, Partial Pressure, Effusion

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SECTION 3 PROBLEMS Write the

answer on the line to the left. Show all

your work in the space provided. 1. 88%

The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the

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percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂ according to the following equation: N₂(g) + 3H₂(g) → 2NH₃(g) N

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Stoichiometry. SECTION 1. SHORT
ANSWER Answer the following questions

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Section 2 Answers

in the space provided. 1. _____ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

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SECTION 9.2. PROBLEMS Write the
answer on the line to the left. Show all
your work in the space provided. 1. The

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Section 2 Answers following equation represents a laboratory preparation for oxygen gas: CHAPTER 9 REVIEW fewer steps are required to solve stoichiometry problems when. ...

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Section 2 Answers

Stoichiometry. SECTION 2. PROBLEMS

Write the answer on the line to the left.

Show all your work in the space provided.

1. The following equation represents a laboratory preparation for oxygen gas:

$2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$... CHAPTER 9 REVIEW ...

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gas produced in part a has density of 1.43
g/L calculate the volume of this gas. 76

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