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Introduction

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Formulas For Ionic

Compounds Writing

Ionic Formulas - Basic

Introduction How To

Name Acids - The

Fast \u0026amp; Easy

Way! How To Name

Ionic Compounds

With Transition Metals

The Criss-Cross

Method for Writing

Chemical Formulas

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~~How to Name~~

~~Chemicals Made~~

~~Easy Naming Ionic~~

~~Compounds with~~

~~Transition Metals~~

~~Introduction Writing~~

~~Formulas with~~

~~Polyatomic Ions How~~

~~to Memorize The~~

~~Polyatomic Ions -~~

~~Formulas, Charges,~~

~~Naming - Chemistry~~

~~*Naming Compounds*~~

~~*in Chemistry* How to~~

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~~calculate valency?~~

VSEPR Theory:

Introduction

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1: Describing

Chemical Reactions

Chemistry Lesson:

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Chemical Formulas

from Compound

Names Pearson

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Chapter 9

Section 5: The Laws

Governing How

Compounds are

Formed *Naming Ionic*

Compounds(HD)

Lewis Diagrams Made

Easy: How to Draw

Lewis Dot Structures

~~Common name,~~

~~chemical name and~~

~~formula of some~~

~~chemical~~

~~compounds.:~~ Valence

Electrons and the

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Periodic Table

*Naming Ionic
Compounds*

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Names \u0026amp; Page

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Compounds* Pearson

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Section 3: Naming
and Writing Formulas

for Molecular

Compounds ~~GCSE~~

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Balancing Chemical
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Ions Empirical
Formula and
Molecular Formula
Introduction *Covalent
Compounds: Writing
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Terms. NaomiSee21.

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Terms 47 Terms.

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20 Terms.

annearnold. OTHER

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PRELIM 194 Terms.

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23 Terms.

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STUDY. PLAY.

Monatomic Ion.

Consists of a single atom with a positive or negative charge resulting from the loss or gain of one or more valence electrons and behaves as a unit, yet it still has a charge. (+ or -). Determined by the amount of electrons lost or gained.

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~~Answers Page
221~~

To name a binary molecular compound, use the following guidelines: 1. Write the names of the elements in the order listed in the formula 2. Use prefixes appropriately to indicate the number of

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each kind of atom 3.

End the name of the second element with the suffix -ide

Answers Page

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...

First, confirm that the compound is a binary molecular compound. Name the elements in the order listed in the

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formula. Use prefixes to indicate the number of each kind of atom. Omit the prefix -mono when the formula contains only one atom of the first element in the name. The suffix of the name of the second element is -ide.

~~Study Chapter 9:~~

~~Chemical Names and~~

Page 17/40

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Formulas Flashcards

Names

1. when the name of the anion (X) ends in -ide the acid name begins with the prefix hydro-. The stem of the anion has the suffix -ic and is followed by the word acid (ex. HCl= hydrochloric acid)
2. when the anion name ends in -ite the acid

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name is the stem of the anion with the suffix -ous followed by the word acid (ex. H_2SO_3 -- X=sulfite = sulfurous acid)

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~~Chemical Names And~~

~~Formulas Test~~

~~Review ...~~

Writing the Names

and Formulas for

Bases. A base is a

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compound that produces OH^- in water. When naming a base, you name it like any other compound that starts with a regular or transition metal. Ex: NaOH = sodium hydroxide. When writing the formula for a base, remember to balance charges. Ex: magnesium hydroxide

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Mg(OH)_2

Names

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Names and

Formulas 83

SECTION 9.3

NAMING AND

WRITING

FORMULAS FOR

MOLECULAR

COMPOUNDS

(pages 268–270) This

Page 22/40

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Section explains the rules for naming and writing formulas for binary molecular compounds. Naming Binary Molecular Compounds (pages 268–269) 1. Circle the letter of the type(s) of elements that form binary molecular compounds.

~~Name Date Class~~

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~~CHEMICAL NAMES~~

~~AND FORMULAS 9~~

~~SECTIONS 9.1~~

~~NAMES AND~~

~~FORMULAS~~

~~FORMULAS FOR~~

~~IONIC COMPOUNDS~~

1. Write the formulas for these binary ionic compounds. c.

potassium iodide 2.

write the formulas for the compounds

formed c. b. f. 3.

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Name the following

binary ionic

compounds. sodium

sulfide g. CuCl : h.

sncl. i. a. MnO_2 c. d.

SrBr_2 221 Chapter 9

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Assessment 42. a. 2–

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b. 1+ c. 1- d. 3+ 43.

a. 2+ b. 2+ c. 3+ d. 1+

44. a. barium ion b.

iodide ion c. silver ion

d. mercury(II) ion 45.

cyanide, CN^- and

hydroxide, OH^- 46. a.

hydroxide ion b.

lead(IV) ion c. sulfate

ion d. oxide ion 47.

zero; A compound is

electrically neutral.

48. The symbols for

the cation and

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~~Guide - Quia~~

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Names and Formulas.

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Gravity. Created by.

bibiloni1901. Terms in

this set (20)

Monatomic Ion.

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Consists of a single atom with a positive or negative charge resulting from the loss of gain of one or more valence

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friend; Find other activities; Start over; Help; Use these activities to help yourself to learn the vocabulary and major topics presented in this chapter. A B; anion: any atom or group of atoms that has a negative charge:

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Page 32/40

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vocabulary for dech's

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Terms in this set (7)
monoatomic ions.

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Vocabulary: Chemical
Names and Formulas

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Names Formulas Test
Answers~~

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~~Answers Page 221 ...~~

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Writing Chemical
Formulas Answer
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Chapter 9

1. ANS: A PTS: 1 DIF:

L1 REF: p. 253 OBJ:

9.1.1 Identify the charges of monatomic ions by using the

periodic table, and

name the ions. STA:

Ch.5.a 2.ANS: H PTS:

1 DIF: L1 REF: p. 254

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~~Test—Naming and~~

~~Writing Chemical~~

~~Formulas~~

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